

S/137/62/000/002/C08/1/
A006/A101

AUTHORS: Bulatova, L. G., Vil'nyanskiy, Ya. Ye.

TITLE: Interaction of chromic chloride with oxygen

PERIODICAL: Referativnyy zhurnal, Metallurgiya, no. 2, 1962, 18, abstract 2A85
("Metallurg. i khim. prom-st' Kazakhstana. Nauchno-tekhn. sb.",
1961, no. 2(12), 68-73)

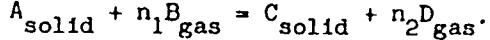
TEXT: Preliminary thermodynamical calculations of chromic and chromous chloride reactions with O_2 revealed the possibility in principle of the full transformation of Cr chlorides into Cr_2O_3 . The reaction $2CrCl_3(\text{solid}) + 1.5 O_2(\text{gas}) = Cr_2O_3(\text{solid}) + 3Cl_2(\text{gas})$ was experimentally investigated. The rate of this reaction was determined as a function of O_2 concentration, the velocity of the gas flow, temperature and duration of interaction. Experimental data are represented in graphs in $\lg K$ versus $(1/T) \cdot 10^3$ coordinates. The nature of graphs and the activation energy values E show that at $350 - 400^\circ\text{C}$, the process takes place in the kinetic range ($E = 31-34 \text{ kcal/mole}$). At $500 - 600^\circ\text{C}$ diffusion conditions take place ($E = 7 - 10 \text{ kcal/mole}$); in the $400 - 500^\circ\text{C}$ range there is an intermediate zone; ($E = 19 \text{ kcal/mole}$). Experimental data presented in

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"transformation degree versus time" coordinates, are S-shaped curves. The average rate, calculated by dividing the degree of CrCl_3 transformation into Cr_2O_3 by time, was taken as the reaction rate. On the basis of the curve type the conclusion is drawn that this reaction is of an autocatalytic nature. At the beginning of the process the reaction rate depends considerably on the velocity of the air current. From 15 minute and more duration, the reaction rate depends very weakly on the gas flow velocity. At low O_2 concentrations and short duration of the reaction, the rate of the latter depends on O_2 concentration in the gas according to an equation of the first order. At 1% O_2 in the initial gas, the rate of CrCl_3 transformation into Cr_2O_3 is $0.5 - 0.7\text{%/min}$ at 500°C . This entails the practical conclusion on the necessity of protecting hot CrCl_3 , obtained during chlorination of ore or Fe-Cr, against the effect of air O_2 during cooling. The interaction of CrCl_3 with O_2 belongs to type 3 of topochemical reactions (according to Roginskiy) and can be schematically described by equation



G. Frents

[Abstracter's note: Complete translation]

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S/598/61/000/005/004/010
D040/D113

AUTHORS: Bezukladnikov, A.B., and Vil'nyanskiy, Ya.Ye.

TITLE: The kinetics of titanium dioxide chlorination in molten chlorides

SOURCE: Akademicheskaya nauk SSSR. Institut metallurgii. Titan i ego splavy, no. 5, Moscow, 1961. Metallurgiya i khimiya titana, 135-142

TEXT: The purpose of the described laboratory experiments was to study the effect of chlorine and oxygen concentration, and of ferrous and aluminum chlorides forming in the chlorination process of titanium-containing raw materials in a medium of molten chlorides; this effect was not determined hitherto. The initial materials used were: carnallite prepared from pure potassium and magnesium chloride obtained by hot magnesium reduction of $TiCl_4$, aluminum chloride; ground petroleum coke; $FeCl_2$ and $FeCl_3$ produced by chlorinating metallic iron in carnallite; TiO_2 boiled for two hours in hydrochloric acid and rinsed in distilled water to wash out chlorine ions; commercial chlorine and argon. Undiluted chlorine, chlorine diluted with ar-

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Gas and air was blown through the melt in the reaction vessel. Observations made at different temperature ranges, proved that the dependence of the chlorination rate on temperature can be described by the Arrhenius equation. No considerable change in the chlorination rate was stated in blowing chlorine diluted by argon to 40% concentration, but further dilution to 30% slowed the chlorination to a half of the rate. Dilution of chlorine by air to oxygen content above 10% speeded up the reaction between $TiCl_4$ and O_2 , but reduced the speed of the chlorination process. Increasing content of $FeCl_3$ or $AlCl_3$ in the melt speeded up the chlorination in all the studied conditions and these chlorides proved to be catalysts. A detailed description is also given of experiment techniques and calculations of equilibrium reaction conditions. Conclusions: (1) The rates of petroleum coke burning and of the TiO_2 chlorination process in molten carnallite depend exponentially on temperature and are within the kinetic field; (2) Chlorine concentration in gas mixture lowered below 40% slows down the chlorination rate; (3) Oxygen content in gas mixture above 10% slows down the chlorination process; (4) Additions of ferrous and aluminum chlorides raise the

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chlorination rate several times; (5) At 2% by weight of FeCl_3 in molten carnallite, the TiO_2 chlorination process passes over from the kinetic to the diffusion field at 600°C . There are 4 figures and 3 tables.

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BEZUKLADNIKOV, A.B.; VIL'NYANSKIY, Ya.Ye.

Effect of iron and aluminum chlorides on the rate of chlorination
of titanium dioxide. Zhur. prikl. khim. 34 no.1:49-53 Ja '61.

(MIRA 14:1)

1. Bereznikovskiy filial Vsesoyuznogo alyuminiyev-magniyevogo
instituta i Ural'skiy politekhnicheskiy institut.
(Iron chloride) (Aluminum chloride)
(Titanium oxide)

S/020/61/140/003/016/020
B103/B101

AUTHORS: Shchegrov, L. N., and Vil'nyanskiy, Ya. Ye.

TITLE: Process and products of hydrolytic decomposition of titanium tetrachloride

PERIODICAL: Akademiya nauk SSSR. Doklady, v. 140, no. 3, 1961, 620-622

TEXT: The interaction between $TiCl_4$ vapor and water vapor was studied in a dynamic system between 25 and $750^{\circ}C$ at molar ratios of $TiCl_4:H_2O = 1:1; 1:2; 1:3; \text{ and } 1:4$. A stream of inert carrier gas was saturated with a definite quantity of $TiCl_4$ vapor, and another one with H_2O . Then, the two gas streams were united in a reactor at test temperature. Contrary to R. F. Hudson (Ref. 1, see below), a white aerosol was formed immediately when the two gas streams met. The degree of hydrolytic decomposition of $TiCl_4$ is not noticeably affected by the average time (11.85 - 0.23 sec) for which the reagents are kept in the reaction zone. Consequently, the hydrolysis of $TiCl_4$ in the vapor phase sets in very rapidly. The particles of solid reaction products formed between 25 and $150^{\circ}C$ are very fine, but are

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agglutinated to form shapeless hygroscopic lumps, so that their further investigation has to be performed in a dry air chamber. Their yellow color gets paler and paler with increasing hydrolyzation temperature and molar concentration of the water in the reaction mixture. The products recovered are good dielectrics. Their composition is a function of the production conditions. A change of the $TiCl_4:H_2O$ ratio from 1:1 to 1:3 results in a reduction of the content of Ti and Cl in the products of hydrolysis. At a ratio of 1:4, the Ti content increases again, whereas the Cl content decreases in products obtained at 25, 50, and 75°C. This is related to the interaction between the products of hydrolysis and the vapor of "excess" water effecting the liberation of HCl. The composition of the products of hydrolysis is described by the formulas: $TiO_a(OH)_{2-2a}Cl_2$, where $a = 0 - 1$; $Ti(OH)_xCl_{(4-x)} \cdot nH_2O$, where $x:(4-x) = 1 - 3$. Thus, the hydrolytic decomposition of $TiCl_4$ cannot be defined by a simple equation. Even at 300°C, hydrolytic decomposition of $TiCl_4$ was not completed (15% of Cl ions in the reaction product). At 500°C, TiO_2 was formed with 2% Cl ions at

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most. At 750°C, the content of Cl ions was 0.1-0.2%. The graph representing the relation between the increase in weight of the hydrolytic products obtained between 25 and 100°C and the time for which they were kept in air begins with a steep rise, attains a maximum after about 1 hr (ca 34% increase in weight), followed by a slight fall. This is ascribed to a chemical reaction between $TiCl_4$ and H_2O , in which water adsorption is accompanied by liberation of HCl. After some time the rate of water adsorption equals the rate of hydrolysis (peak of the graph). Subsequently, hydrolysis prevails over sorption. By using Debye-Scherrer patterns it has been found that the products of $TiCl_4$ hydrolysis in the vapor phase between 75 and 150°C form a new crystalline phase which is not found any more at 300°C. Between 25 and 50°C and at a ratio of $TiCl_4 : H_2O = 1:1, 1:2,$ and $1:3$, x-ray diffraction analysis reveals that the products have no crystalline structure. After storage in moist air, their structure becomes crystalline like that of the products obtained at 75 and 150°C. Between 300 and 750°C, TiO_2 of anatase structure is formed. There are 3 figures and 6 references: ✓
2 Soviet and 4 non-Soviet. The three references to English-language publications read as follows: Ref. 1: R. F. Hudson, Proc. of the XI Intern. Card 3/4

Process and products of...

S/020/61/140/003/016/020
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Congr. of Pure and Appl. Chem., 1, London, 1947, p. 297; Ref. 2: R. L.
Powell, Chem. Eng. Progr., 50, 11, 578 (1954); Ref. 3: L. W. Rowe, W. R.
Opie, J. Metals, 7 (11), sect. 1, 1189 (1955).

ASSOCIATION: Ural'skiy politekhnicheskiy institut im. S. M. Kirova
(Ural Polytechnical Institute imeni S. M. Kirov)

PRESENTED: April 14, 1961 by S. I. Vol'fkovich, Academician
SUBMITTED: April 2, 1961

Card 4/4

SAVINKOVA, Ye.I.; ORSKHOVA, A.I.; VIL'NYANSKIY, Ya.Ye.; YANKOVSKIY, V.R.

Quick determination of phase composition in synthetic cinnallite.
TSvet. met. 38 no.9:58-60 S '65.

(MIRA 18:12)

"APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9

SHCHEGROV, I.N.; VIL'NYANSKIY, Ya.Ye.; BAYBAKOV, D.P.

Synthesis of the products of titanium tetrachloride hydrolysis.
Trudy IREA no.25:470-478 '63.
(MIRA 18:6)

APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9"

SOLOSHENKO, A.A.; VIL'NYANSKIY, Ya.Ye.

Kinetics of hydrogen chloride oxidation on a chromium oxide catalyst. Kin. i kat. 5 no.5:881-887 S-O '64. (MIRA 17:12)

1. Ural'skiy nauchno-issledovatel'skiy khimicheskiy institut.

SAVINKOVA, Ye.I.; BYCHIKHINA, L.S.; VIL'NYANSKIY, Ya.Ye.

Effect of the composition of carnallite on its hydrolysis in
the atmosphere of hydrogen chloride and water vapor. Zhur.
prikl. khim. 37 no.6:1356-1358 Je '64.

1. Ural'skiy politekhnicheskiy institut imeni Kirova. (MIRA 18:3)

KROTOV, Yu.I.; VIL'NYANSKIY, Ya.Ye.; KOGAN, B.S.

Function of a nickel electrode as oxygen electrode in
oxygen-containing salts. Zhur. fiz. khim. 38 no.6:1632-
1635 Je '64.
(MIRA 18:3)

1. Ural'skiy politekhnicheskiy institut imeni Kirova.

"APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9

UKHANOVA, T.L.; VIL'NIYANERIY, Ya.Ye.

Solubility of hydrogen chloride in chloride melts. Izv. vys. ucheb. zav.; khim. i khim. tekhn. 7 no.3:510-513 '64.

1. Ural'skiy politekhnicheskiy institut imeni Kirova, kafedra tekhnologii neorganicheskikh veshchestv. (MIRA 17:10)

APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9"

TETEREVKOV, A.I.; VIL'NYANSKIY, Ya.Ye.

Mass transfer rate in the chlorination of a magnesium oxide
suspension in molten chlorides. Izv.vys.ucheb.zav.;khim.i khim.tekh.
6 no.5:874-878 '63. (MIRA 16:12)

1. Ural'skiy politekhnicheskiy institut imeni Kirova, kafedra
tekhnologii neorganicheskikh veshchestv.

VIL'NYANSKIY, Ya.Ye.; BOROVSKIKH, L.A.; KOVEL', M.S.

Preparation of chromium oxide by reducing alkaline chromate
with sulfur dioxide. Zhur.VKHO 8 no.1:116-117 '63.

1. Ural'skiy politekhnicheskiy institut imeni S.M. Kirova.
(Chromium oxides) (Sodium chromate) (Sulfur dioxide)

(MIRA 16:4)

RYABIN, V.A.; VIL'NYANSKIY, Ya.Ye.; PAVLOVA, V.P.

Certain variations in the phase composition of the reaction mass
in the process of oxidative calcination of chromite charges.
Dokl.AN SSSR 149 no.3:652-655 Mr '63. (MIRA 16:4)

1. Ural'skiy politekhnicheskiy institut im. S.M.Kirova.
Predstavleno akademikom S.I.Vol'fkovichem.
(Chromates) (Oxidation)

IURZHENKO, A.I.; VIL'SHANSKIY, V.A.

Emulsion polymerization with surface-active initiation. Dokl.
AN SSSR 148 no.5:1145-1147 F '63. (MIRA 16:3)

1. Predstavлено академиком P.A.Rebinderom.
(Polymerization) (Surface-active agents)

SAVINKOVA, Ye.I.; BOROVSKIKH, L.A.; VIL'NYANSKIY, Ya.Ye.

Relation between the speed of chlorination of the magnesium oxide suspension in chloride melt and the nature of the reducing agent and its dosage. Trudy Ural. politekh. inst. no.94:48-52 '60.

(MIRA 15:6)

(Magnesia) (Chlorination)

BEZUKLADNIKOV, A.B.; VIL'NYANSKIY, Ya.Ye.

Kinetics of chlorinating titanium dioxide in a media of fused
chlorides. Titan i ego splavy no.5:135-142 '61. (MIRA 15:2)
(Chlorination)
(Titanium oxide)

S/137/62/000/004/030/201
A006/A101

AUTHORS: Devyatovskaya, L. I., Vil'nyanskiy, Ya. Ye.

TITLE: Preparation of chromium chlorides from ferrochromium

PERIODICAL: Referativnyy zhurnal, Metallurgiya, no. 4, 1962, 28, abstract 4G180
("Metallurg. i khim. prom-st' Kazakhstana. Nauchno tekhn. sb.",
1961, no. 3 (13), 59-64)

TEXT: Batches of 50 - 70 g Fe-Cr were processed with Cl-gas in a vertical laboratory furnace at 950 - 1,000°C. Fully sublimated CrCl₃, CrCl₂ and FeCl₃ were obtained. Separation of Cr and Fe from the condensate was not checked but, according to literature data, satisfactory results can be obtained. Chlorination of Fe-Cr proceeds at a higher rate and with 98% utilization of Cl-gas.

A. Tseydler

[Abstracter's note: Complete translation]

Card 1/1

SHCHEGROV, I.N.; VIL'NYANSKIY, Ya.Ye.

Process and products of the hydrolytic decomposition of titanium tetrachloride. Dokl. Akad. Nauk SSSR 140 no.3:620-622 S '61. (MIRA 14:9)

1. Ural'skiy politekhnicheskiy institut im. S.M.Kirova. Predstavлено
академиком S.I.Vol'fkovichem.
(Titanium chloride) (Hydrolysis)

S/137/62/000/005/042/150
A006/A101

AUTHORS: Teterevkov, A. I., Vil'nyanskiy, Ya. Ye.

TITLE: Preparation of chromium chloride by ferrochrome chlorination with hydrogen chloride

PERIODICAL: Roferativnyy zhurnal, Metallurgiya, no. 5, 1962, 25, abstract 50157 ("Metallurg. i khim. prom-st' Kazakhstana. Nauchno-tekhn. sb.", 1961, no. 5(15), 58 - 65)

TEXT: The authors studied interaction of HCl with Fe-Cr at 600 - 1,000°C. Chlorination of Fe-Cr was conducted in boats and vertical retorts. It was established that the interaction rate of HCl with Fe-Cr increases with temperature, and Cr and Fe dichlorides are formed as a result of reactions. With a higher rate of HCl supply at 1,000°C, the process rate rises and HCl utilization decreases. Molten CrCl₂ with ≤ 1% FeCl₂ can be obtained by the action of HCl upon Fe-Cr at 950 - 1,000°C. The process can be conducted in a shaft-furnace-type apparatus with counterflow of gas and Fe-Cr. An increase in the Fe-Cr surface raises the HCl utilization and the chlorination rate. G. Svoitseva

[Abstracter's note: Complete translation]
Card 1/1

S/137/62/000/006/031/163
ACC6/A101

AUTHORS: Bezukladnikov, A. B., Vil'nyanskiy, Ya. Ye.

TITLE: Kinetics of chlorinating titanium dioxide in molten chloride medium

PERIODICAL: Referativnyy zhurnal, Metallurgiya, no. 6, 1962, 13, abstract 6097
(In collection: "Titan i yego splavy", no. 5, Moscow AN SSSR, 1961,
135 - 142)

TEXT: The authors studied the nature of the effect of $FeCl_3$ and $AlCl_3$ forming during the chlorination process, and of Cl and O concentration upon the chlorination rate of TiO_2 in molten carnallite. Chlorination with 100% Cl was conducted at 500 - 900°C. The process runs in the kinetic range. Chlorination of Cl, with diluted argon was conducted at 900°C. A reduction of Cl concentration in a gas mixture to < 40% reduces considerably the chlorination rate. During the chlorination of TiO_2 with a Cl and O mixture an increase of the O content to > 10% accelerates the rate of $TiCl_4$ interaction with O and reduces the chlorination rate. Admixtures of Fe and Al chlorides accelerate the chlorination rate several times; this is explained by their catalytic effect. In

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Kinetics of...

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the presence of 2% FeCl₃ in the carnallite melt, TiO₂ chlorination passes over from the kinetic to the diffusion range at 680°C.

L. Vorob'yeva

[Abstracter's note: Complete translation]

Card 2/2

BYCHIKHINA, L.S.; VIL'NYANSKIY, Ya.Ye.; SAVINKOVA, Yo.I.

Gasometric method of determining the hydrogen content in products
of the partial hydrolysis of carnallite and magnesium chloride,
Izv. vys. ucheb. zav.; tsvet. met. 5 no.2:100-103 '62.
(MIRA 15:3)

1. Ural'skiy politekhnicheskiy institut, kafedra tekhnologii
neorganicheskikh veshchestv.
(Metals--Hydrogen content) (Magnesium--Metallurgy)

"APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9

VIL'NYY, V.[Vil'nyi, V.]; ZIRKO, M., red.; KONTAR, K., tekhn. red.

[Chernovtsy; photosketches] Chernivtsi; fotonarys. Kyiv, Derzh.
vyd-vo obrazotvorchogo mystetstva i muzychnoi lit-ry, 1962. 6 p.
(MIRA 15:11)

(Chernovtsy—Views)

APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9"

"APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9

VHO, L. (Tallinn)

Conference on engineering geology in Estonia. (En., Russ. & Eng.).
grau. 6 no.5,29-30 '64. (USTRIC 17:17)

APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9"

VILOV, V.

Georgi Dimitrov Water-Power Electric Plant. p. 1.

REKHNIKA. Vol. 4, no. 5, June/July 1955

Sofia, Bulgaria

SOURCE: East European Accessions List (EEAL) Library of Congress, Vol. 6, No. 1, January 1957

"APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9

VILOVIC, Vasja, ing.

Juvinil-granulate.II. Kem ind 10 no.9:259-262,264 S '61.

APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9"

VILOVIC, Vasja, ing.

Juvinil-granulate, polyvinyl chlorides. Kem ind 10 no.6:Supp.J-23
J-26, J-28 Je '61.

VIIKOVAYA, K. and AKHALONVS, T.

Condition of the eyes after protracted application of plasmocid
Vestnik Oftalmologii 1940, 17/5 (630-632)

5219 In 532 malaria patients, treated exclusively with plasmocid and plasmocid combined with acrichin during a period of two years, atrophy of the optic nerve was observed in only one case. No data are given in the article to prove that the atrophy developed in consequence of plasmocid-poisoning.

Kalfus- Odessa (See XII)

SO: Section II Vol. 1² No. 7-12

"APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9

VILOVSKI, Slavomir; VLAHINIC-DEKIC, Ksenija

Geology of the Gola Glava Mountain to the northeast of Jajce
(Western Bosnia). Geol glas BiH 7:69-80 '63.

APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9"

"APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9

VILOVSKI, Slavomir; LAUSEVIC, Miroslav

Geologic structure of the Plasa Mountain, Hercegovina. Geol
glas BiH 9:77-85 '64.

1. Submitted May 30, 1964.

APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9"

VILOVSKIY, N.S.

In the fight to raise labor productivity. Elek. i tepl. tiaga 3
no.2:13 F '59. (MIRA 12:4)

1. Starshiy master zagotovitel'nogo tsentral'nogo depo Likhobory.
(Railroads—Repair shops)

VILOVSKIY, A.L.; KAZANSKIY, V.I.

Coronary Arteries - Diseases

Development of collateral cardiac circulation in coronary disorders. Klin. med., 30,
no. 1, 1952.

Monthly List of Russian Accessions, Library of Congress, May 1952, UNCLASSIFIED

VIL'PERT, A.

River regulation in the Northern Basin. Rech. transp. 22
no.9:43-45 S '63. (MIRA 16:10)

1. Nachal'nik sluzhby puti Severnogo basseynovogo upravleniya
puti.

"APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9

VIL'PKRT, A.S., inzhener.

More on expenditures per one kilometer of waterway. Rech.transp. 14
no.12:21-23 Je '55.
(Waterways) (Dredging) (MIRA 9:9)

APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9"

DAYCHIK, M.L., kand.tekhn.nauk; VIL'FERT, K.I.; VORONKOV, V.A.

Devices for statistical investigations of accelerations, stresses,
and deformations. Avt.prom. 29 no.10:22-25 O '63. (MIRA 16:10)

1. Gosudarstvennyy soyuznyy ordena Trudovogo Krasnogo Znameni
nauchno-issledovatel'skiy avtomobil'nyy i avtomotornyy institut.

NIKOLAYEV, V.I.; VIL'PERT, K.I.

Device for measuring and recording instantaneous values of the ignition-advance angle and the number of revolutions of an engine. Avt.prom. 28 no.5:26-27 My '62. (MIRA 15:5)

1. Gosudarstvennyy soyuзnyy ordena Trudovogo Krasnogo Znameni nauchno-issledovatel'skiy avtomobil'nyy i avtomotornyy institut.
(Motor vehicles--Engines--Testing)
(Electronic instruments)

"APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9

VIL'PERT, K.I.

Semiconductor d.c. converters. Avt. prom. 30 no.6:46-47
Je '64.

(MIRA 17:12)

1. TSentral'nyy ordena Trudovogo Krasnogo Znameni nauchno-
issledovatel'skiy avtomobil'nyy i avtomotornyy institut.

APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9"

VIL'PERT, K.I.; PEVZNER, Ya.M., doktor tekhn.nauk; TIKHONOV, A.A., kand.tekhn.
nauk; YUDIN, B.V.

Some problems in the statistical analysis of vibrations of a
motor vehicle. Avt.prom. 31 no.4:26-29 Ap '65.

1. TSentral'nyy ordena Trudovogo Krasnogo Znameni nauchno-issledo-
vatel'skiy avtomobil'nyy i avtomotornyy institut. (MIRA 18:5)

"APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9

BAI CIUNAS, Jonas; DOVYDAITIS, Vytautas; OSKINIS, Bronius; VILPISauskas, V.,
red.; PRONSKIETYTE, D., red.; PAKERYTE, O., tekhn. red.

[Following the clouds] Debesu keliais. Vilnius, Valstybine
politines ir mokslynes literaturos leidykla, 1961. 190 p.
(MIRA 15:3)
(Lithuania--Gliding and soaring)

APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001859820009-9"

JASIUKEVICIUS, V.; JARULAITIS, V.; LASYS, A.; SASNAUSKAS, K.;
ZUBAUSKAS, A.; VILPISAUSKAS, V., red.; MONTRIMAS, R.,
red.; CECYTE, V., tekh. red.

[Production of bricks, tiles, and drainpipes] Plytu cerpiu ir
dremu gamyba. [By] V.Jasiukevicius ir kiti. Vilnius, Valstybine
politines ir mokslyines literaturo leidykla, 1961. 258 p.

(Bricks) (Tiles) (Drain tiles) (MIRA 15:3)

DILBA, I., inzh.; KRIKSTOPAITIS, I., inzh.; KVILIUS, L., inzh.;
RASIULIS, B., inzh.; SIDARAVICIUS, L., inzh.; STRIMAITIS, C.,
inzh.; VILPISAUSKAS, V., red.; KUOSAITE, R., red.; PAKENYTE, O.,
tekhn. red.

[A concise builder's guide] Trumpas statybininko vadovas.
[By] I.Dilba ir kiti. Vilnius, Valstybine politines ir
mokslines literaturos leidykla, 1961. 395 p.

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